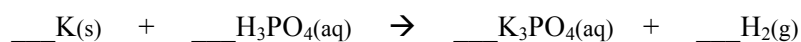


### AE Solution Stoichiometry

In the space below copy the entire Stoichiometry Flow Chart from your notes to use as a reference as you do the problems on this worksheet.

Use the following reaction to answer the questions below.



1. When 13.0 g of K react, what concentration of  $\text{K}_3\text{PO}_4$  will result if the final solution volume is 2.0 L?
2. If 0.300 L of 2.0 M  $\text{H}_3\text{PO}_4$  are used, what volume of **hydrogen gas** will be produce at STP?
3. If 300 mL of 0.25 M  $\text{K}_3\text{PO}_4$  are produced, what mass of K was used?
4. When 44.8 L of hydrogen gas were produced, what was the initial concentration of  $\text{H}_3\text{PO}_4$  used?  
Assume the volume of solution was 1.5 L.